Title: Focus Chemistry Form 5 2021

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Errata

Page number	Section/ Part	Error	Correction				
Page 32	Comparison between Electrolytic Cell and Voltaic Cell	Voltaic cell Electrolytic cell Chemical cell Chemical reactions occur at the electrodes producing electric current Electrical energy → chemical energy Conversion of energy Electrical energy → electrical energy Electrical energy → electrical energy	Differences Voltaic cell				
Page 52	Checkpoint	S4 S5 S6	Q4 Q5 Q6				
Page 56	2 (Figure 2.7)	Akhiran Name it appropriately as meth-, eth-, prop-, but-, pent-, etc (Table 2.5)	Suffix Name it appropriately as meth-, eth-, prop-, but-, pent-, etc (Table 2.7)				
Page 65	Table 2.15	5 th column: Solubility (g per 100 g of water) Bercampur dalam semua kadaran Bercampur dalam semua kadaran Bercampur dalam semua kadaran	5 th column: Solubility (g per 100 g of water) Miscible in all proportions Miscible in all proportions Miscible in all proportions				
Page 70	Figure 2.15	Misplaced of 'Alkene' with 'Alkane'	Exchange label of 'Alkene' with 'Alkane' or vice versa				
Page 97	Table 2.21	2 nd column: Molecular formula, C _n H _{2n-2} 4 5	2^{nd} column: Molecular formula, C_nH_{2n-2} C_4H_6 C_5H_8				

	Mixture Sodium hydroxide + water	(°C)	(°C)	or absorbed	Type of reaction	Mixture	Initial temperature (°C)	Highest/ lowest temperature	Heat is released or absorbed	Type of reaction
	Anhydrous calcium chloride + water	33	41			Sodium		(C)		
Results Figure 3.9	Ammonium chloride + water	33	29				33	39	Heat is released	Exothermic
	Sodium thiosulphate + water	33	31	Heat is absorbed	Exothermic	water				
					,	Anhydrous calcium chloride + water	33	41	Heat is released	Exothermic
						Ammonium chloride + water	33	29	Heat is absorbed	Endothermic
						Sodium thiosulphate + water	33	31	Heat is absorbed	Endothermic
			ΔH = positive	nic reaction			<u> </u>	ΔH = po	psitive	1
Solution	Number of moles of HNO ₃ = $\frac{MV}{1000}$ = $\frac{M \times 26}{1000}$ = 0.026 mol			Number of moles of HNO ₃ = $\frac{MV}{1000}$ $= \frac{M \times 26}{1000}$						
	Figure 3.9	Figure 3.9 Figure 3.9 Figure 3.9 Figure 3.9 Figure 3.9 Figure 3.	Mixture temperature (°C) Sodium hydroxide + water 33 Anhydrous calcium chloride + water 33 Ammonium chloride + water 33 Sodium thiosulphate + water 33 Results Energy Figure 3.9 Products Figure 3.9 Energy level diagramment Figure 3.9 Energy level diagramment Number of moles of HNO3 = $\frac{MV}{1000}$ $\frac{MV}{1000}$ Solution $\frac{M \times 26}{1000}$	Results Mixture temperature temperature (°C) Sodium hydroxide + water 33 39 Anhydrous calcium chloride + water 33 41 Ammonium chloride + water 33 29 Sodium thiosulphate + water 33 31 Results Energy Reactants Products Figure 3.9 Energy level diagram for endothern Number of moles of HNO ₃ = $\frac{MV}{1000}$ Solution Solution $\frac{MV}{1000}$ $\frac{MV}{10000}$ M	Results C	Results Continue Continue	Results Sodium hydroxide + water 33 39 Heat is released Exothermic Anhydrous calcium chloride + water 33 41 Heat is released Exothermic Nydroxide + water Sodium thiosulphate + water 33 31 Heat is absorbed Exothermic Nydroxide + water Sodium So	Results Colum hydroxide + water 33 39 Heat is released Exothemic Sodium hydroxide + water 33 29 Heat is absorbed Exothemic Sodium hydroxide + water 33 31 Heat is paloribed Exothemic Sodium hydroxide + water 33 31 Heat is absorbed Exothemic Sodium hydroxide + water 33 31 Heat is absorbed Exothemic Sodium hydroxide + water 33 31 Heat is absorbed Exothemic Ammonium chloride + water 33 31 Heat is absorbed Exothemic Ammonium Ammoniu	Mixture temperature (PCO) codium hydroxide + water 33 39 Heat is released Exothemic (PC) Sodium hydroxide + water 33 39 Heat is absorbed Exothemic (PC) Sodium hydroxide + water 33 29 Heat is absorbed Exothemic (PC) Sodium hydroxide + water 33 39 Heat is absorbed Exothemic (PC) Sodium hydroxide + water 33 39 Heat is absorbed Exothemic (PC) Sodium hydroxide + water 33 41 Heat is absorbed Exothemic (PC) Sodium hydroxide + water 33 41 Results Products Energy Energy Energy Energy Figure 3.9 Energy level diagram for endothermic reaction Figure 3.9 Energy level diagram for endothermic reaction Number of moles of HNO ₃ = MV MV 26	Results Mixture temperature temperatu